

## Balancing Equations Worksheet

- 1) \_\_\_\_  $\text{Na}_3\text{PO}_4$  + \_\_\_\_  $\text{KOH}$   $\rightarrow$  \_\_\_\_  $\text{NaOH}$  + \_\_\_\_  $\text{K}_3\text{PO}_4$
- 2) \_\_\_\_  $\text{MgF}_2$  + \_\_\_\_  $\text{Li}_2\text{CO}_3$   $\rightarrow$  \_\_\_\_  $\text{MgCO}_3$  + \_\_\_\_  $\text{LiF}$
- 3) \_\_\_\_  $\text{P}_4$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{P}_2\text{O}_5$
- 4) \_\_\_\_  $\text{RbNO}_3$  + \_\_\_\_  $\text{BeF}_2$   $\rightarrow$  \_\_\_\_  $\text{Be}(\text{NO}_3)_2$  + \_\_\_\_  $\text{RbF}$
- 5) \_\_\_\_  $\text{AgNO}_3$  + \_\_\_\_  $\text{Cu}$   $\rightarrow$  \_\_\_\_  $\text{Cu}(\text{NO}_3)_2$  + \_\_\_\_  $\text{Ag}$
- 6) \_\_\_\_  $\text{CF}_4$  + \_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_  $\text{CBr}_4$  + \_\_\_\_  $\text{F}_2$
- 7) \_\_\_\_  $\text{HCN}$  + \_\_\_\_  $\text{CuSO}_4$   $\rightarrow$  \_\_\_\_  $\text{H}_2\text{SO}_4$  + \_\_\_\_  $\text{Cu}(\text{CN})_2$
- 8) \_\_\_\_  $\text{GaF}_3$  + \_\_\_\_  $\text{Cs}$   $\rightarrow$  \_\_\_\_  $\text{CsF}$  + \_\_\_\_  $\text{Ga}$
- 9) \_\_\_\_  $\text{BaS}$  + \_\_\_\_  $\text{PtF}_2$   $\rightarrow$  \_\_\_\_  $\text{BaF}_2$  + \_\_\_\_  $\text{PtS}$
- 10) \_\_\_\_  $\text{N}_2$  + \_\_\_\_  $\text{H}_2$   $\rightarrow$  \_\_\_\_  $\text{NH}_3$
- 11) \_\_\_\_  $\text{NaF}$  + \_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_  $\text{NaBr}$  + \_\_\_\_  $\text{F}_2$
- 12) \_\_\_\_  $\text{Pb}(\text{OH})_2$  + \_\_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_\_  $\text{H}_2\text{O}$  + \_\_\_\_  $\text{PbCl}_2$
- 13) \_\_\_\_  $\text{AlBr}_3$  + \_\_\_\_  $\text{K}_2\text{SO}_4$   $\rightarrow$  \_\_\_\_  $\text{KBr}$  + \_\_\_\_  $\text{Al}_2(\text{SO}_4)_3$
- 14) \_\_\_\_  $\text{CH}_4$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CO}_2$  + \_\_\_\_  $\text{H}_2\text{O}$
- 15) \_\_\_\_  $\text{Na}_3\text{PO}_4$  + \_\_\_\_  $\text{CaCl}_2$   $\rightarrow$  \_\_\_\_  $\text{NaCl}$  + \_\_\_\_  $\text{Ca}_3(\text{PO}_4)_2$
- 16) \_\_\_\_  $\text{K}$  + \_\_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_\_  $\text{KCl}$
- 17) \_\_\_\_  $\text{Al}$  + \_\_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_\_  $\text{H}_2$  + \_\_\_\_  $\text{AlCl}_3$
- 18) \_\_\_\_  $\text{N}_2$  + \_\_\_\_  $\text{F}_2$   $\rightarrow$  \_\_\_\_  $\text{NF}_3$
- 19) \_\_\_\_  $\text{SO}_2$  + \_\_\_\_  $\text{Li}_2\text{Se}$   $\rightarrow$  \_\_\_\_  $\text{SSe}_2$  + \_\_\_\_  $\text{Li}_2\text{O}$
- 20) \_\_\_\_  $\text{NH}_3$  + \_\_\_\_  $\text{H}_2\text{SO}_4$   $\rightarrow$  \_\_\_\_  $(\text{NH}_4)_2\text{SO}_4$

## Balancing Equations Worksheet – Answers

**Note to students:** It is acceptable to leave spaces blank when balancing equations – blank spaces are interpreted as containing the number "1".

- 1) 1  $\text{Na}_3\text{PO}_4$  + 3  $\text{KOH}$   $\rightarrow$  3  $\text{NaOH}$  + 1  $\text{K}_3\text{PO}_4$
- 2) 1  $\text{MgF}_2$  + 1  $\text{Li}_2\text{CO}_3$   $\rightarrow$  1  $\text{MgCO}_3$  + 2  $\text{LiF}$
- 3) 1  $\text{P}_4$  + 3  $\text{O}_2$   $\rightarrow$  2  $\text{P}_2\text{O}_5$

# 33 Physical Science Balancing Worksheet 2 Answers

**Carol Lillis, Priscilla LeMone, Marilee  
LeBon, Pamela Lynn**



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**General Catalogue of Printed Books** British Museum. Department of Printed Books, 1969

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## **Table of Contents 33 Physical Science Balancing Worksheet 2 Answers**

1. Understanding the eBook 33 Physical Science Balancing Worksheet 2 Answers
  - The Rise of Digital Reading 33 Physical Science Balancing Worksheet 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 33 Physical Science Balancing Worksheet 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 33 Physical Science Balancing Worksheet 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 33 Physical Science Balancing Worksheet 2 Answers
  - Personalized Recommendations
  - 33 Physical Science Balancing Worksheet 2 Answers User Reviews and Ratings
  - 33 Physical Science Balancing Worksheet 2 Answers and Bestseller Lists
5. Accessing 33 Physical Science Balancing Worksheet 2 Answers Free and Paid eBooks

- 33 Physical Science Balancing Worksheet 2 Answers Public Domain eBooks
  - 33 Physical Science Balancing Worksheet 2 Answers eBook Subscription Services
  - 33 Physical Science Balancing Worksheet 2 Answers Budget-Friendly Options
6. Navigating 33 Physical Science Balancing Worksheet 2 Answers eBook Formats
- ePub, PDF, MOBI, and More
  - 33 Physical Science Balancing Worksheet 2 Answers Compatibility with Devices
  - 33 Physical Science Balancing Worksheet 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 33 Physical Science Balancing Worksheet 2 Answers
  - Highlighting and Note-Taking 33 Physical Science Balancing Worksheet 2 Answers
  - Interactive Elements 33 Physical Science Balancing Worksheet 2 Answers
8. Staying Engaged with 33 Physical Science Balancing Worksheet 2 Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 33 Physical Science Balancing Worksheet 2 Answers
9. Balancing eBooks and Physical Books 33 Physical Science Balancing Worksheet 2 Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection 33 Physical Science Balancing Worksheet 2 Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 33 Physical Science Balancing Worksheet 2 Answers
- Setting Reading Goals 33 Physical Science Balancing Worksheet 2 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 33 Physical Science Balancing Worksheet 2 Answers
- Fact-Checking eBook Content of 33 Physical Science Balancing Worksheet 2 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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