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# **Chapter 7 Chemical Reactions**

# Summary

#### 7.1 Describing Reactions

- In a chemical reaction, the substances that undergo change are called reactants.
- The new substances formed as a result of that change are called products.
- A chemical equation is a representation of a chemical reaction in which the reactants and products are expressed as formulas.

The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction.

 During a chemical reaction, the mass of the products is always equal to the mass of the reactants.

In order to show that mass is conserved during a reaction, a chemical equation must be balanced.

- In a balanced chemical equation, the number of atoms on the left side of the equation equals the number of atoms on the right.
- You can balance a chemical equation by changing the coefficients.
- Coefficients are the numbers that appear before the formulas.

Because chemical reactions often involve large numbers of small particles, chemists use a counting unit called the mole to measure amounts of a substance.

- A mole (mol) is an amount of a substance that contains approximately 6.02 × 10<sup>to</sup> particles of that substance. This number is known as Avogadro's number.
- The mass of one mole of a substance is called a molar mass. For an element, the molar mass is the same as its atomic mass expressed in grams.

In chemical reactions, the mass of a reactant or product can be calculated by using a balanced chemical equation and molar masses of the reactants and products.

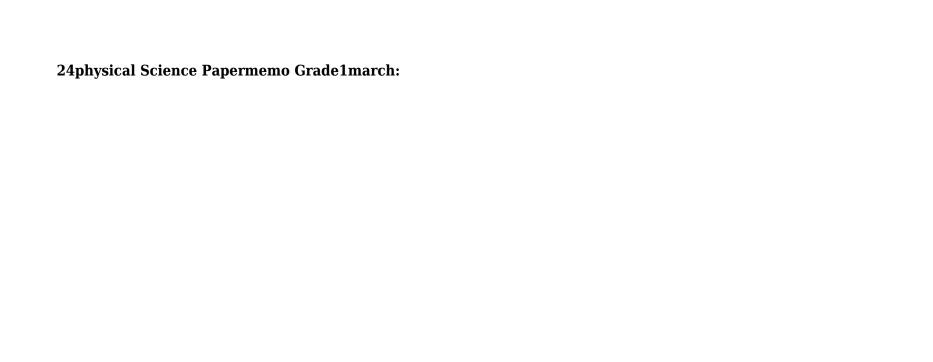
#### 7.2 Types of Reactions

Some general types of chemical reactions are synthesis reactions, decomposition reactions, single-replacement reactions, doublereplacement reactions, and combustion reactions.

 A synthesis reaction is a reaction in which two or more substances react to form a single substance. The general equation for a synthesis reaction is A + B → AB.

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