

## Tutorial 1: Chemical Kinetics Answers

1. Consider the following concentration-time data for the reaction of iodide ion and hypochlorite ion ( $\text{OCl}^-$ ). The products are chloride ion and hypiodite ion ( $\text{OI}^-$ ).

| Experiment | Time (s) | $[\text{I}^-]$        | $[\text{OCl}^-]$      | $[\text{OH}^-]$ |
|------------|----------|-----------------------|-----------------------|-----------------|
| 1          | 0        | $2.40 \times 10^{-4}$ | $1.60 \times 10^{-2}$ | 1.00            |
|            | 10       | $2.16 \times 10^{-4}$ | $1.36 \times 10^{-2}$ | 1.00            |
| 2          | 0        | $1.20 \times 10^{-4}$ | $1.60 \times 10^{-2}$ | 1.00            |
|            | 10       | $1.08 \times 10^{-4}$ | $1.48 \times 10^{-2}$ | 1.00            |
| 3          | 0        | $2.40 \times 10^{-4}$ | $4.00 \times 10^{-3}$ | 1.00            |
|            | 10       | $2.34 \times 10^{-4}$ | $3.40 \times 10^{-3}$ | 1.00            |
| 4          | 0        | $1.20 \times 10^{-4}$ | $1.60 \times 10^{-2}$ | 2.00            |
|            | 10       | $1.14 \times 10^{-4}$ | $1.54 \times 10^{-2}$ | 2.00            |

- (a) Write a balanced equation for the reaction.  
 (b) Determine the rate law, and calculate the value of the rate constant.  
 (c) Does the reaction occur by a single-step mechanism? Explain.  
 (d) Propose a mechanism that is consistent with the rate law, and express the rate constant in terms of the rate constants for elementary steps in your mechanism. (Hint: Transfer of an  $\text{H}^+$  ion between  $\text{H}_2\text{O}$  and  $\text{OCl}^-$  is a rapid reversible reaction.)



- b) assume that  $\text{rate} = \frac{-\Delta[\text{I}^-]}{\Delta t} = k[\text{OCl}^-]^a[\text{I}^-]^b[\text{OH}^-]^c$   
 compare experiments 1 and 3:

$$\frac{2.4 \times 10^{-4} - 2.17 \times 10^{-4}}{2.4 \times 10^{-4} - 2.34 \times 10^{-4}} = \left( \frac{1.6 \times 10^{-2}}{4 \times 10^{-3}} \right)^a$$

$$4^a \approx 4$$

$$\therefore a = 1$$

compare experiments 1 and 2:

$$\frac{2.4 \times 10^{-4} - 2.17 \times 10^{-4}}{1.2 \times 10^{-4} - 1.08 \times 10^{-4}} = \left( \frac{2.4}{1.2} \right)^b$$

$$2^b \approx 2$$

$$\therefore b = 1$$

compare experiments 1 and 4:

$$\frac{2.4 \times 10^{-4} - 2.17 \times 10^{-4}}{1.2 \times 10^{-4} - 1.14 \times 10^{-4}} = 2 \left( \frac{1}{2} \right)^c$$

$$2 \left( \frac{1}{2} \right)^c \approx 4$$

$$\therefore c = -1$$

$$\text{rate} = \frac{-\Delta[\text{I}^-]}{\Delta t} = k \frac{[\text{OCl}^-][\text{I}^-]}{[\text{OH}^-]}$$

$$\frac{1.2 \times 10^{-4} - 1.08 \times 10^{-4}}{10} = k \frac{1.2 \times 10^{-4}(1.6 \times 10^{-2})}{1}$$

# 1 Review And Reinforcement Chemical Kinetics Answers

**S Ben Porath**



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## **Table of Contents 1 Review And Reinforcement Chemical Kinetics Answers**

1. Understanding the eBook 1 Review And Reinforcement Chemical Kinetics Answers
  - The Rise of Digital Reading 1 Review And Reinforcement Chemical Kinetics Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 1 Review And Reinforcement Chemical Kinetics Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 1 Review And Reinforcement Chemical Kinetics Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 1 Review And Reinforcement Chemical Kinetics Answers
  - Personalized Recommendations
  - 1 Review And Reinforcement Chemical Kinetics Answers User Reviews and Ratings
  - 1 Review And Reinforcement Chemical Kinetics Answers and Bestseller Lists
5. Accessing 1 Review And Reinforcement Chemical Kinetics Answers Free and Paid eBooks
  - 1 Review And Reinforcement Chemical Kinetics Answers Public Domain eBooks
  - 1 Review And Reinforcement Chemical Kinetics Answers eBook Subscription Services
  - 1 Review And Reinforcement Chemical Kinetics Answers Budget-Friendly Options
6. Navigating 1 Review And Reinforcement Chemical Kinetics Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - 1 Review And Reinforcement Chemical Kinetics Answers Compatibility with Devices
  - 1 Review And Reinforcement Chemical Kinetics Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 1 Review And Reinforcement Chemical Kinetics Answers
  - Highlighting and Note-Taking 1 Review And Reinforcement Chemical Kinetics Answers
  - Interactive Elements 1 Review And Reinforcement Chemical Kinetics Answers
8. Staying Engaged with 1 Review And Reinforcement Chemical Kinetics Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 1 Review And Reinforcement Chemical Kinetics Answers
- 9. Balancing eBooks and Physical Books 1 Review And Reinforcement Chemical Kinetics Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 1 Review And Reinforcement Chemical Kinetics Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 1 Review And Reinforcement Chemical Kinetics Answers
  - Setting Reading Goals 1 Review And Reinforcement Chemical Kinetics Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1 Review And Reinforcement Chemical Kinetics Answers
  - Fact-Checking eBook Content of 1 Review And Reinforcement Chemical Kinetics Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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