

SAPLING CHEMISTRY ANSWERS SAMPLE

Question: Calculate the number of Carbon Dioxide (CO_2) molecules in a 2.44×10^{-2} mole of CO_2 .

Solution:

$N = n \cdot N_A$ Avogadro constant (N_A)

= (number of particles, typically molecules or atoms contained in the amount of substance provided by a single mole.

$N_A = 6.02 \times 10^{23} \text{ mole}^{-1}$

Final Answer:

$N = 2.44 \times 10^{-2} (6.02 \times 10^{23}) = 1.47 \times 10^{22} \text{ molecules}$

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JG Myers



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