

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

- Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

| | | | | |
|-----------|-----------------------------------------------------------|------|------|-------------------------------------|
| Equation: | $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ | | | 1.2 moles KClO_3 needed |
| Before: | 1.2 | 0 | 0 | |
| Change: | -1.2 | +1.2 | +1.8 | |
| After: | 0 | 1.2 | 1.8 | |

- In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

| | | | | | |
|-----------|-------------------------------------------------------------------|----------|---------|-------|------------------------------------|
| Equation: | $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ | | | | 0.200 moles HCl needed |
| Before: | 0.100 | ? | 0 | 0 | |
| Change: | -0.100 | -0.200 + | 0.100 + | 0.100 | |
| After: | 0 | ? | 0.100 | 0.100 | |

- Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

| | | | | |
|-----------|------------------------------------------------------------------------------------------|-------|-------|-------|
| Equation: | $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$ | | | |
| Before: | 2.20 | xs | 0 | 0 |
| Change: | -2.20 | -7.70 | +4.40 | +6.60 |
| After: | 0 | xs | 4.40 | 6.60 |

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

- Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

| | | | | |
|-----------|-------------------------------------------------------------------------------------------------|---------|---------|---------|
| Equation: | $\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$ | | | |
| Before: | 0.0040 | XS | 0 | 0 |
| Change: | -0.0040 | -0.0080 | +0.0040 | +0.0080 |
| After: | 0 | XS | 0.0040 | 0.0080 |

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

Answer Sheet Chemistry Unit 8 Worksheet 4

Nathaniel Max Rock

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Abstract Bulletin of the Institute of Paper Chemistry Institute of Paper Chemistry (Appleton, Wis.),1958

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