

Answer Sheet

Name: _____

Date: _____

1. (A) (B) (C) (D)

2. (A) (B) (C) (D)

3. (A) (B) (C) (D)

4. (A) (B) (C) (D)

5. (A) (B) (C) (D)

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49. (A) (B) (C) (D)

50. (A) (B) (C) (D)

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Robert R. Redfield



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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{\ln \left(\frac{N_0}{N} \right)}{\ln 2}$ 125 g 5 what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore $n = \frac{\ln \left(\frac{100}{80} \right)}{\ln 2}$ now using the half life equation in example 1 we have

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