

## Lab #13: Qualitative Analysis of Cations and Anions

### Objectives

1. To understand the rationale and the procedure behind the experiments for various cations and anions.
2. To perform qualitative analysis of two unknown solutions that contain various ions (cations and anions) and positively identify them by using established schemes.

### Materials

#### Unknown Cation Solution

Unknown cation (one or all of the following)

- 0.01M Ag nitrate solution,  $\text{AgNO}_3(aq)$
- 0.01M Ni sulphate (II) solution,  $\text{NiSO}_4(aq)$
- 0.01M Al nitrate solution,  $\text{Al(NO}_3)_3(aq)$
- 0.01M Pb acetate (II) solution,  $\text{Pb(CH}_3\text{COO)}_2(aq)$

- 0.1M  $\text{HCl}(aq)$
- 0.1M  $\text{HNO}_3(aq)$
- 0.1M  $\text{NH}_3(aq)$
- 0.1M  $\text{H}_2\text{S}(aq)$
- 0.1M  $\text{NaOH}(aq)$
- 0.1M  $\text{H}_2\text{SO}_4(aq)$
- 0.1M  $\text{Na}_2\text{CO}_3(aq)$
- 0.1M potassium permanganate solution,  $\text{KMnO}_4(aq)$
- 0.1M  $\text{AgNO}_3(aq)$
- Standard  $\text{NaOH}(aq)$
- 0.1M  $\text{BaCl}_2(aq)$
- Adjustment granules

#### Unknown Anion Solution

Unknown anion (one or all of the following)

- 0.01M sodium chloride,  $\text{NaCl}(aq)$
- 0.01M sodium sulphate,  $\text{Na}_2\text{SO}_4(aq)$
- 0.01M sodium nitrate,  $\text{NaNO}_3(aq)$
- 0.01M sodium acetate,  $\text{NaCH}_3\text{COO}(aq)$

- 1.4 Test Tubes and 4 Test Tube Racks
- 400 mL Beaker (Wash)
- 250 mL Beaker
- Spatula
- 4 Marsh-type Pipets
- Hot Plate
- Stirring Rods
- 4 Small Test Tube Shoppers
- Labels or pH Paper
- Wash Glass
- Centrifuge

### Background Information

Qualitative analysis is an analytical procedure in which the question “what is present?” is answered. It is systematic qualitative analysis scheme, such sufficiently present is expected. Since the other substances, there is a confirmatory test is used to prove that the isolated substance is the expected one.

To begin the lab experiment, a solution containing five cations is prepared using the techniques for qualitative analysis of cations. A second solution containing five anions is then prepared using the qualitative scheme for anions. These solutions, one containing any combination of five different cations, and another containing any combination of five anions will be assigned to each student group. The two “unknown” solutions are then analyzed to determine which ions are present and which are absent. This experiment is carried out as a semi-analytic study. Very small quantities of reagents are used. Cleanliness and a great deal of care are necessary to obtain good results. While going through the steps of the analysis, keep a copy of the appropriate flow chart available for reference. The flow charts appear as a later page are simplified versions of what you have learned from the textbook. They will help to give the “total picture” of where each analysis is and where it is heading. Read the directions carefully, and read the paragraphs that give the names or chemical theory for each step. Don’t just follow directions “cook, book” style, but make an effort to understand the chemical principles behind the procedures.

### Heating Substances

Importantly it will be necessary to heat a substance to speed up a reaction. The NEXT lab will use before using Bunsen burner flame. A modern means holds will cause the solution to shoot out of the test tube. Instead, heat test tube in a boiling water bath. A good idea is to set up this bath when beginning work in this lab because it may take time to heat the bath to the appropriate temperature.

# Analysis Of Anions Cations Lab Answers

**Donnell R. Christian, Jr.**



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*Chemistry in the Laboratory* James M. Postma,Julian L. Robert,J. Leland Hollenberg,2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations **Report of Investigations**

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*Annual Report of the Agricultural Experiment Station, Michigan State University* Michigan State University. Agricultural Experiment Station,1919     *Director's Report* Kansas Agricultural Experiment Station,1922     Open File Report ,1983

*Research Bulletin* University of Missouri--Columbia. Agricultural Experiment Station,1925     Energy Research Abstracts ,1990     *Annual Report* New Jersey Agricultural Experiment Station,1923 Includes report of the New Jersey Agricultural College Experiment Station     **CRREL Report** ,1993

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