

Please follow the instructions for each section of the exam. Show your work on all mathematical problems. Provide answers with the correct units and significant figures. Be concise in your answers to discussion questions.

Part 8: Warmup. 4 points each

1. Reactions with a positive ΔH° and a negative ΔS° are:
a. spontaneous at all temperatures.
b. non-spontaneous at all temperatures.
c. spontaneous at low temperatures but non-spontaneous at high temperatures. Answer **B**
d. non-spontaneous at low temperatures but spontaneous at high temperatures.
2. Which of the processes below DOES NOT result in an increase in entropy?
a. $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
b. $\text{H}_2\text{O}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l})$ (The melting of ice.)
c. $\text{CO}_2(\text{s}) \rightarrow \text{CO}_2(\text{g})$ (The sublimation of dry ice.) Answer **A**
d. $\text{NH}_4\text{NO}_3(\text{s}) \rightarrow \text{N}_2\text{O}(\text{g}) + 2 \text{H}_2\text{O}(\text{g})$
3. Consider the reaction below. If hydrogen chloride gas is being produced at 1.2 moles per liter per minute (M min^{-1}), at what rate is hydrogen gas being consumed?
$$2\text{HCl}(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{I}_2(\text{g}) + 2\text{HCl}(\text{g})$$

a. 1.2 M min^{-1}
b. 0.60 M min^{-1}
c. 2.4 M min^{-1}
d. Not enough information to determine. Answer **B**

Part 6: Complete all of problems 4-9

4. Concisely discuss the validity of each of the following statements: a maximum of three sentences each. Clearly justify your reasoning. (12 points)
a. Reactions with a positive ΔH° and a positive ΔS° are always spontaneous.

Your answer should discuss how ΔH° , ΔS° and ΔG° are related. Since $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$, as long as $T\Delta S^\circ$ is more positive than ΔH° , the reaction will be spontaneous. This occurs if ΔS° is very large or at increased temperature.

- b. Free energy changes provide a good indication of which reactions are favorable and fast, as well as those that are unfavorable and slow.

Free energy changes are thermodynamic quantities that tell us about the tendency for a reaction to proceed, they tell us nothing about the rate of the reaction. We need information about kinetics to do that.

Analytical Chemistry Final Exam Multiple Choice

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145 Things to Be When You Grow Up Jodi Weiss, Princeton Review (Firm), 2004-05-04 Like some teenagers you may already have an inkling of what you want to do with your life Perhaps you dream of launching into space as an astronaut with NASA Or maybe your ambitions will keep your feet on the ground as a park ranger a doctor or a social worker If you are wondering what these people do all day and how they came to be what they are this book provides the answers We ll tell you everything you need to know about 145 great careers from accountant to zoologist including Day to day activities and responsibilities of each career How to plan ahead while you re still in high school Majors to consider when you get to college Whether you ll strike it rich in more ways than one Words of wisdom from professionals in the field This book will open your eyes to a world of career possibilities and give you an idea of what to expect down each career path Even if you have no clue about what you want to be this is the book for you We ll help you identify the kinds of jobs that will help you be yourself and incorporate your interests values and skills There s no pressure for you to choose a career just yet so just relax and have fun exploring

America's Best Value Colleges Eric Owens, Princeton Review (Firm), 2004-03-15 This informative guide profiles 77 schools that not only charge less in tuition but are more likely to help students with financial aid scholarships and grants

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