



## Hydrogen Peroxide Analysis Consumer Chemistry

### Introduction

Hydrogen peroxide is regarded as an "environmentally friendly" alternative to chlorine for water purification and wastewater treatment. Because hydrogen peroxide decomposes in the presence of heat, light, or other catalysts, the quality of a hydrogen peroxide solution must be checked regularly to ensure its effectiveness. The concentration of hydrogen peroxide can be analyzed by redox titration with potassium permanganate.

### Concepts

- Redox reaction
- Oxidizing and reducing agents
- Titration
- Half-reactions

### Background

Titration is a method of volumetric analysis—the use of volume measurements to analyze the concentration of an unknown. The most common types of titrations are acid–base titrations, in which a solution of an acid, for example, is analyzed by measuring the amount of a standard base solution required to neutralize a known amount of the acid. A similar principle applies to redox titrations. If a solution contains a substance that can be oxidized, then the concentration of that substance can be analyzed by titrating it with a standard solution of a strong oxidizing agent.

The equation for an oxidation–reduction reaction can be balanced by assuming that it occurs via two separate half-reactions. In this experiment, potassium permanganate will be used as the titrant to analyze the concentration of hydrogen peroxide in a commercial antiseptic solution. The permanganate ion acts as an oxidizing agent—it causes the oxidation of hydrogen peroxide. The oxidation half-reaction shows that two electrons are lost per molecule of hydrogen peroxide that is oxidized to oxygen gas (Equation 1). The permanganate ion, in turn, is reduced from the +7 oxidation state in  $\text{MnO}_4^-$  to the +2 oxidation state in  $\text{Mn}^{2+}$ . The reduction half-reaction shows a gain of five electrons (Equation 2).



### Experiment Overview

The purpose of this experiment is to analyze the percent hydrogen peroxide in a common "drugstore" solution by titrating it with potassium permanganate. Standard potassium permanganate solution will be added via buret to the hydrogen peroxide solution. As the dark purple solution is added, it will react with the hydrogen peroxide and the color will clear. When all of the hydrogen peroxide has been used up, the endpoint of the titration occurs when the last drop of potassium permanganate that is added causes the solution to stay pink.

### Pre-Lab Questions (Answer questions on a separate sheet of paper.)

1. Combine the oxidation and reduction half-reactions for hydrogen peroxide and permanganate ion, respectively, and write the balanced chemical equation for the overall reaction between  $\text{H}_2\text{O}_2$  and  $\text{MnO}_4^-$  in acid solution. *Hint:* The number of electrons transferred must "cancel out."
2. What is the mole ratio of hydrogen peroxide to permanganate ion in the balanced chemical equation determined in Question #1? How many moles of hydrogen peroxide will be oxidized by 0.0045 moles of potassium permanganate in acidic solution?

# Analysis Of Hydrogen Peroxide Lab Answers

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