

# Two-Step Equation Word Problems

Write and solve an equation for each problem.



- 1 Carla is a waitress at Daybreak Diner, and she earns \$5 for each hour she works. Last week, she earned \$148 total, including \$68 in tips. How many hours did Carla work last week?

- 2 Mr. Olin bought 6 sketchbooks for the students in his art class. Each sketchbook was on sale for \$3 off. If the sketchbooks cost \$18 total before tax, what was the original price of each sketchbook?

- 3 Davion wants to buy 4 tickets to a basketball game. He sees that there is a discount for \$25 off his entire purchase at checkout, so the tickets would cost him \$215 before tax. What was the original price of each ticket?

- 4 Elana makes knit hats and scarves and sells them in her online store for \$12 each. Last week, Elana's store made \$180 total. If she sold 8 hats, how many scarves did she sell?

- 5 Tiana bought 3 video games that each cost the same amount. The tax on each video game was \$1.29. She paid a total of \$68.37 after tax. What was the price of each video game, before tax?

- 6 The local pool near Paxton's house charges a one-time registration fee of \$18.25 and \$17.50 per month. Last year, Paxton paid \$105.75. How many months was Paxton a member of the pool?

# 2 Step Equation Word Problems With Answers

**Reza Nazari, Ava Ross**



## **2 Step Equation Word Problems With Answers:**

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web 1 practice problem balance the redox reaction in basic solution  $\text{NO}_2^{-} + \text{Al} + \text{S} + \text{NH}_3 \rightarrow \text{AlO}_2^{-} + \text{N}_2 + \text{S}^{2-}$  practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of  $\text{OH}^{-}$   $\text{Br}_2 + \text{I}^{-} \rightarrow \text{Br}^{-} + \text{I}_2$  previous topic next topic

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web google classroom for the following redox equation  $\text{MnO}_4^{-} + \text{SO}_2 + 2\text{H}^{+} \rightarrow \text{Mn}^{2+} + \text{H}_2\text{SO}_4$  what will be the stoichiometric coefficient of  $\text{H}_2\text{SO}_4$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} \rightarrow \text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} + 2\text{OH}^{-} \rightarrow 2\text{Au}(\text{CN})_2^{-} + 2\text{H}_2\text{O} + \text{O}_2$  multiplied by a factor of 2  $2\text{H}_2\text{O} + \text{O}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{OH}^{-}$  3 the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution  $4\text{Fe}^{2+} + 6\text{OH}^- + 7\text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{NO}_3^- + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox

reaction in acid solution using the half reaction method 8 h<sub>2</sub>o<sub>2</sub> cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> 2 9 teo<sub>3</sub> 2 n<sub>2</sub>o<sub>4</sub> o<sub>2</sub> cr<sub>3</sub> te 10 reo<sub>4</sub> 11 pbo<sub>2</sub> io<sub>2</sub> pb<sub>2</sub> 12 as no<sub>3</sub> io<sub>3</sub> re io<sub>3</sub> h<sub>2</sub>aso<sub>4</sub>

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h<sup>+</sup> 2no<sup>2-</sup> 2 c 6 h 6 o 6 2h<sup>+</sup> 2 o<sup>2-</sup> 2no choose 1 answer c 6 h 8 o 6 and no<sup>2-</sup> a c 6 h 8 o 6 and no<sup>2-</sup> no 2 and c 6 h 8 o 6 b no<sup>2-</sup> and c 6 h 8 o 6 h<sup>+</sup> and no<sup>2-</sup> c h<sup>+</sup> and

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