

ACIDS, BASES, & THE pH SCALE

Name: **Answer Key**

Date: _____

Pd: _____

PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
<ul style="list-style-type: none">• Sour taste, leaves burning sensation• Increase hydrogen ions (H^+) in solution• Example: HCl $H^+ + Cl^-$• Turns litmus (pH) paper to red	<ul style="list-style-type: none">• Sharp bitter taste, feels slippery to touch• Decrease hydrogen ions (H^+) in solution (Increase hydroxide ions (OH^-) in solution)• Example: NaOH $Na^+ + OH^-$• Turns litmus (pH) paper to blue

PART 2: THE pH SCALE

- The **pH Scale** is a **measurement** of how acidic or how basic a solution is.
- The scale goes from **0** to **14**.
- **Compounds** are **acidic** if they have a pH **lower** than **7**.
- Compounds with a pH **higher** than **7** are **basic** or **alkaline**.
- Compounds with a pH **of 7**, are **neutral**.

0 1 2 3 4 5 6 7 8 9 10 11 12 13 14

Increased concentration of **H^+** ions
Decreased concentration of **OH^-** ions

Decreased concentration of **H^+** ions
Increased concentration of **OH^-** ions

pH below 7 = **acids**

pH at 7 = **neutral**
(**pure water**)

pH above 7 = **bases**

PART 3: BUFFERS

- A **buffer** is a **weak acid** or **base** that **prevents** sudden **changes** in pH.
 - Maintains **homeostasis**
- A buffer works by **binding** to H^+ ions **releasing** H^+ ions when the solution pH changes.
- Buffers **bind** to H^+ ions when the **concentration** of H^+ ions in solution **increases**.
- **NaOH** **$Na^+ + OH^-$** (The OH^- will bind to H^+ ions to form H_2O – **neutralization**)
- Buffers **release** H^+ ions when the **concentration** of H^+ ions in solution **decreases**.
HCl **$H^+ + Cl^-$**

20 Acids And Bases Answer Key

Disha Experts

A red circular graphic with a gradient, appearing as a partial circle or a thick curved line, located to the right of the 'Disha Experts' text.

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General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value

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