

ACIDS, BASES, & THE pH SCALE

Name: **Answer Key**

Date: _____

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PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
• Sour taste, leaves burning sensation • Increase hydrogen ions (H^+) in solution • Example: HCl $H^+ + Cl^-$ • Turns litmus (pH) paper to red	• Sharp bitter taste, feels slippery to touch • Decrease hydrogen ions (H^+) in solution (Increase hydroxide ions (OH^-) in solution) • Example: NaOH $Na^+ + OH^-$ • Turns litmus (pH) paper to blue

PART 2: THE pH SCALE

- The **pH Scale** is a **measurement** of how acidic or how basic a solution is.
- The scale goes from **0** to **14**.
- **Compounds** are **acidic** if they have a pH **lower** than **7**.
- Compounds with a pH **higher** than **7** are **basic** or **alkaline**.
- Compounds with a pH **of 7**, are **neutral**.

0 1 2 3 4 5 6 7 8 9 10 11 12 13 14

Increased concentration of **H^+** ions
Decreased concentration of **OH^-** ions

Decreased concentration of **H^+** ions
Increased concentration of **OH^-** ions

pH below 7 = **acids**

pH at 7 = **neutral**
(**pure water**)

pH above 7 = **bases**

PART 3: BUFFERS

- A **buffer** is a **weak acid** or **base** that **prevents** sudden **changes** in pH.
 - Maintains **homeostasis**
- A buffer works by **binding** to H^+ ions **releasing** H^+ ions when the solution pH changes.
- Buffers **bind** to H^+ ions when the **concentration** of H^+ ions in solution **increases**.
- **NaOH** **$Na^+ + OH^-$** (The OH^- will bind to H^+ ions to form H_2O – **neutralization**)
- Buffers **release** H^+ ions when the **concentration** of H^+ ions in solution **decreases**.
HCl **$H^+ + Cl^-$**

20 1 Describing Acids And Bases Answer Key

Lars I. Eriksson



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