

Name: \_\_\_\_\_ ANSWER KEY \_\_\_\_\_ Class: \_\_\_\_\_ Date: \_\_\_\_\_

### Unique Properties of Water!

#### Word Bank:

Adhesion	Evaporation	Polar	Surface tension
Cohesion	Freezing	Positive	Universal solvent
Condensation	Melting	Sublimation	
Dissolve	Negative		

1. The electrons are not shared equally between the hydrogen and oxygen atoms of water creating a Polar molecule.
2. The polarity of water allows it to dissolve most substances. Because of this it is referred to as the universal solvent.
3. Water molecules stick to other water molecules. This property is called cohesion.
4. Hydrogen bonds form between adjacent water molecules because the positive charged hydrogen end of one water molecule attracts the negative charged oxygen end of another water molecule.
5. Water molecules stick to other materials due to its polar nature. This property is called adhesion.
6. Hydrogen bonds hold water molecules closely together which causes water to have high surface tension. This is why water tends to clump together to form drops rather than spread out into a thin film.
7. Condensation is when water changes from a gas to a liquid.
8. Sublimation is when water changes from a solid directly to a gas.
9. Freezing is when water changes from a liquid to a solid.
10. Melting is when water changes from a solid to a liquid.
11. Evaporation is when water changes from a liquid to a gas.
12. Why does ice float?  
Water expands as it freezes, so it is LESS DENSE AS A SOLID.
13. What property refers to water molecules resembling magnets? How are these alike?  
Polar bonds create positive and negative ends of the molecule. The positive ends are attracted to other negative ends, creating hydrogen bonds between molecules.

## 2 Properties Of Water Answers

**Jennifer E. Lawson, Rosalind  
Poon, Deidre Sagert, Melanie  
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