

# NATURAL SCIENCES

## Gr. 9 - LESSON PLAN - MATTER & MATERIALS

TOPIC: COMPOUNDS		
Sub-topic: Names of Compounds		Duration: 1 Hour
CONTENT & CONCEPTS (CAPS p63)		
<b>KEY CONCEPTS</b> <ol style="list-style-type: none"> <li>1. Chemical reaction</li> <li>2. Chemical formula</li> <li>3. Names of compounds according to their elements.</li> </ol>		
Specific Aims:	Specific Aim 1: 'Doing Science'	X
	Specific Aim 2: 'Knowing the subject content and making connections'	X
	Specific Aim 3: 'Understanding the uses of Science'	
<b>LESSON OBJECTIVES:</b> <ol style="list-style-type: none"> <li>1. Learners will be able to write the names of compounds as indicated on page 63 of the CAPS Document.</li> <li>2. Learners will be able to make models of different compounds</li> </ol>		
<b>RESOURCES REQUIRED:</b> <ul style="list-style-type: none"> <li>• Periodic Table Poster</li> <li>• Siyavula Sasol Inzalo (Learner book) page 163</li> </ul>		
<b>TEACHING &amp; LEARNING ACTIVITY:</b> <b>Introduction</b> <u>Briefly review and revise concepts dealt with in Grade 8, focusing on compounds</u>  As each element has a unique name for example: <ul style="list-style-type: none"> <li>- Hydrogen</li> <li>- Lithium</li> </ul> Remember the first 20 elements learnt in grade 8. How did you memorize them? Do you remember Mnemonics... <b>He</b> - H; <b>He</b> - He; <b>Lies</b> - Li; <b>Because</b> - Be; <b>Boys</b> - B; <b>Can</b> - C; <b>Not</b> - N; <b>Operate</b> - O; <b>Fireplaces</b> - F;  <b>New</b> - Ne; <b>Nation</b> - Na; <b>Might</b> - Mg; <b>Also</b> - Al; <b>Sign</b> - Si; <b>Peace</b> - P; <b>Security</b> - S; <b>Clause</b> - Cl  <b>A</b> - Ar; <b>King</b> - K; <b>Can</b> - Ca LEARNERS MAY RECITE THIS MNEMONIC AS A WAY OF MEMORIZING THE FIRST 20 ELEMENTS.  <u>The teacher provides the following explanation to the learners</u>  Each <b>compound</b> has a unique name. How is a compound formed? <ul style="list-style-type: none"> <li>- A compound is formed when TWO or more different elements combine chemically.</li> <li>- Each atom in a compound are held by chemical bonds</li> <li>- Compounds are formed during <b>chemical reaction (when elements/ compounds combine to form new substances)</b></li> <li>- Each compound can be written as a <b>formula</b>, for example, <b>H<sub>2</sub>O</b>; <b>NaCl</b>; <b>CuSO<sub>4</sub></b></li> </ul> What is a <b>formula</b> ? <ul style="list-style-type: none"> <li>- This is a notation that shows which elements the compound is made up of.</li> <li>- It indicates the total number of atoms in one unit of the compound</li> <li>- It indicates the ratio in which atoms of different elements are bonded to each other</li> <li>- It gives the total number of atoms in one unit of a compound.</li> </ul> For example, the formula for Copper Sulphate is CuSO <sub>4</sub>		

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**RJ Shavelson**



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## **Table of Contents 2question Paper Natural Science G**

1. Understanding the eBook 2question Paper Natural Science G
  - The Rise of Digital Reading 2question Paper Natural Science G
  - Advantages of eBooks Over Traditional Books
2. Identifying 2question Paper Natural Science G
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2question Paper Natural Science G
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2question Paper Natural Science G
  - Personalized Recommendations
  - 2question Paper Natural Science G User Reviews and Ratings
  - 2question Paper Natural Science G and Bestseller Lists
5. Accessing 2question Paper Natural Science G Free and Paid eBooks
  - 2question Paper Natural Science G Public Domain eBooks
  - 2question Paper Natural Science G eBook Subscription Services
  - 2question Paper Natural Science G Budget-Friendly Options
6. Navigating 2question Paper Natural Science G eBook Formats
  - ePub, PDF, MOBI, and More
  - 2question Paper Natural Science G Compatibility with Devices
  - 2question Paper Natural Science G Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 2question Paper Natural Science G
  - Highlighting and Note-Taking 2question Paper Natural Science G
  - Interactive Elements 2question Paper Natural Science G

8. Staying Engaged with 2question Paper Natural Science G
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2question Paper Natural Science G
9. Balancing eBooks and Physical Books 2question Paper Natural Science G
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2question Paper Natural Science G
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 2question Paper Natural Science G
  - Setting Reading Goals 2question Paper Natural Science G
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2question Paper Natural Science G
  - Fact-Checking eBook Content of 2question Paper Natural Science G
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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