

▼ Part A

If a reaction mixture initially contains a CO concentration of 0.1510 M and a Cl_2 concentration of 0.175 M at 1000 K . What is the equilibrium concentration of CO at 1000 K ?

Express your answer in molarity to three significant figures.

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$\text{I} \square \square$

$?$

$[\text{CO}] =$ M

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✖ Incorrect; Try Again; 2 attempts remaining

It appears that you have not included the equilibrium constant in your calculation. Recall that the equilibrium concentration of CO is related to the equilibrium constant and the change in concentrations of the product and reactants:

$$K_c = \frac{[\text{COCl}_2]}{[\text{CO}][\text{Cl}_2]}$$

You may want to review [Hint 1. Prepare an ICE table.](#)

▼ Part B

What is the equilibrium concentration of Cl_2 at 1000 K ?

Express your answer in molarity to three significant figures.

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$[\text{Cl}_2] =$ M

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Modern Chemistry Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001

A Laboratory Manual of General Chemistry for Use in Colleges William Crowell Bray,Wendell Mitchell Latimer,1921

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