

# Gas Variables

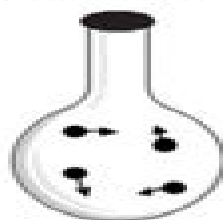
How are the variables that describe a gas related?

## Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar<sup>®</sup>, which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

## Model 1 – Gases in a Nonflexible Container

### Experiment A (Adding more gas)



A1

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 1 atm  
Temperature = 200 K



A2

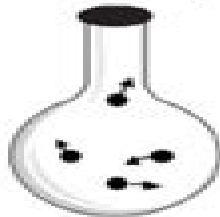
Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 2 atm  
Temperature = 200 K



A3

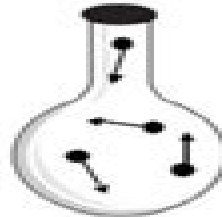
Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 3 atm  
Temperature = 200 K

### Experiment B (Heating the gas)



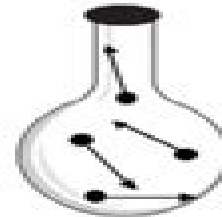
B1

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 1 atm  
Temperature = 200 K



B2

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 2 atm  
Temperature = 400 K



B3

Volume = 1 unit  
External pressure = 1 atm  
Internal pressure = 3 atm  
Temperature = 600 K

\*Note: Volume in this model is recorded in units rather than liters because 4 molecules of gas at the conditions given would occupy a very small space ( $\sim 1 \times 10^{-22}$  mL). The particles shown here are much larger compared to the space between them than actual gas particles.

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