

Planck also proposed the *Quantum Theory* which states that energy is radiated or absorbed in the form of packets or bundles, called quanta (singular, quantum) and each quantum is equal to  $hf$ .

## 2. Photoelectric Effect

Albert Einstein used the quantized energy idea of Planck to come up with the Photon Theory which proposes that light is composed of photons. According to this theory, photon has the following characteristics:

- It has a velocity of light in free space. The velocity of light ( $c$ ) is equal to  $3 \times 10^8$  m/s.
- It has rest energy.
- It has zero mass.
- It can carry energy and momentum.
- When radiation is emitted or absorbed, it can be created or destroyed.
- It can interact with other particles like electrons.

Using his photon theory, Einstein explained the phenomenon called *photoelectric effect*. In this phenomenon, when light strikes a metal surface, the electrons would become excited and jump out of the metal. This is because, according to Einstein, light consists of packets of energy called photons which transfer energy to the electron causing it to become excited.

## 3. Atomic Spectra

In your Grade 9 science, you learned that electrons orbits or energy levels. As long as electron stays in the orbit, it does not absorb or emit energy and said to be in its ground state. But when the atom is being heated, its electrons become excited and occupy a higher energy level. At this point the electron is in its excited state. As electrons leave the excited state and return to its ground state again, they emit energy in the form of photon.

The set of frequencies of electromagnetic spectrum emitted by excited atom is referred to as the atomic emission spectra. Every element emits a specific energy and wavelength producing a unique set of banding pattern, thus we can say that atomic spectra serves as fingerprints of every element. The figure below shows the emission spectrum of an electrically excited hydrogen atom. The fine lines corresponds to the emitted light with a specific wavelength while the dark bands in between corresponds to the absorbed frequencies.



Figure1. Emission spectrum of hydrogen atom

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