#### ANSWERS

#### Nuclear Chemistry Review Questions

Compare the mass, charge, and penetration power of alpha, beta and gamma radiation.

	Mass	Charge	Penetration
Alpha	4	+	Lowest
Beta	0	-	Middle
Gamma	0	0	Highest

In 5.49 seconds, 1.20g of argon-35 decays to leave only 0.15g. What is the half-life?

1.20g → 0.60g → 0.30g → 0.15g , therefore it took 3 half-life periods.

5.49 + 3 = 1.83 sec

3. How many days does it take 16.0g of gold-198 to decay to 1.0g?

16.0g → 8.0g → 4.0g → 2.0g → 1.0g , therefore 4 half-life periods

From Ref Table N, 1 half-life = 2.69d 2.69d \* 4 = 10.76 days

4. Why does an atom undergo radioactive decay?

Because it has an unstable nucleus

5. Compare and contrast fission and fusion. Why is fission used for electrical and fusion is not?

Fission=Division, the splitting of large unstable atoms(nucleus)
Fusion=joining together, small isotopes of H, creates way more energy

6. How is nuclear chemistry applied to medicine? What type of radioisotope must be used when it is injected into a person?

Radioisotopes are used to treat and diagnose illnesses
If injected they must have short half-lives and quickly removed from the body

7. Write out the nuclear decay of radium-226.

236 Ra→ He+ 222 Rn

8. What is different about nuclear reactions when compared to a chemical reaction?

Nuclear reactions transform atoms into other atoms, chemical reactions involved the rearrangement of electrons

At what point do all isotopes of an element become radioactive?
 Up to and including atomic # 83

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**Hussin A.Rothana** 

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## Table of Contents 25 Nuclear Chemistry Guided Study Answers 236161

- 1. Understanding the eBook 25 Nuclear Chemistry Guided Study Answers 236161
  - The Rise of Digital Reading 25 Nuclear Chemistry Guided Study Answers 236161
  - Advantages of eBooks Over Traditional Books
- 2. Identifying 25 Nuclear Chemistry Guided Study Answers 236161
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 25 Nuclear Chemistry Guided Study Answers 236161
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from 25 Nuclear Chemistry Guided Study Answers 236161
  - Personalized Recommendations
  - o 25 Nuclear Chemistry Guided Study Answers 236161 User Reviews and Ratings
  - 25 Nuclear Chemistry Guided Study Answers 236161 and Bestseller Lists

- 5. Accessing 25 Nuclear Chemistry Guided Study Answers 236161 Free and Paid eBooks
  - o 25 Nuclear Chemistry Guided Study Answers 236161 Public Domain eBooks
  - 25 Nuclear Chemistry Guided Study Answers 236161 eBook Subscription Services
  - 25 Nuclear Chemistry Guided Study Answers 236161 Budget-Friendly Options
- 6. Navigating 25 Nuclear Chemistry Guided Study Answers 236161 eBook Formats
  - o ePub, PDF, MOBI, and More
  - 25 Nuclear Chemistry Guided Study Answers 236161 Compatibility with Devices
  - 25 Nuclear Chemistry Guided Study Answers 236161 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 25 Nuclear Chemistry Guided Study Answers 236161
  - Highlighting and Note-Taking 25 Nuclear Chemistry Guided Study Answers 236161
  - Interactive Elements 25 Nuclear Chemistry Guided Study Answers 236161
- 8. Staying Engaged with 25 Nuclear Chemistry Guided Study Answers 236161
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 25 Nuclear Chemistry Guided Study Answers 236161
- 9. Balancing eBooks and Physical Books 25 Nuclear Chemistry Guided Study Answers 236161
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection 25 Nuclear Chemistry Guided Study Answers 236161
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 25 Nuclear Chemistry Guided Study Answers 236161
  - Setting Reading Goals 25 Nuclear Chemistry Guided Study Answers 236161
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 25 Nuclear Chemistry Guided Study Answers 236161
  - Fact-Checking eBook Content of 25 Nuclear Chemistry Guided Study Answers 236161
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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