

If solubility of  $\text{PbSO}_4 = s \text{ M}$

$$\therefore [\text{Pb}^{2+}] = s$$

$$[\text{SO}_4^{2-}] = s + 1.2 \times 10^{-2} \approx 1.2 \times 10^{-2} \quad (\text{assuming } s \ll 1.2 \times 10^{-2})$$

$$\therefore s \times 1.2 \times 10^{-2} = 1.6 \times 10^{-8}$$

$$s = \frac{1.6}{1.2} \times 10^{-6} = 1.33 \times 10^{-6}$$

On comparing with  $X \times 10^{-Y}$

$$Y = 6$$

2. An aqueous solution is prepared by dissolving 0.1 mol of an ionic salt in 1.8 kg of water at  $35^\circ\text{C}$ . The salt remains 90% dissociated in the solution. The vapour pressure of the solution is 59.724 mm of Hg. Vapor pressure of water at  $35^\circ\text{C}$  is 60.000 mm of Hg. The number of ions present per formula unit of the ionic salt is \_\_\_\_\_.

**Answer (5)**

**Sol.** Number of ions present per formula unit of ionic salt =  $x$

$$\text{Van 't Hoff factor } (i) = 0.9 \times x + 0.1 \times 1$$

(Assuming 90% dissociation)

$$\therefore \text{Relative lowering in vapour pressure} = \text{Mole fraction of solute}$$

$$\Rightarrow \frac{60 - 59.724}{60} = \frac{i \times 0.1}{\frac{1800}{18} + 0.1}$$

$$\Rightarrow 0.0046 = \frac{i \times 0.1}{100 + 0.1}$$

$$0.0046 = \frac{(0.9x + 0.1) \times 0.1}{100}$$

$$\Rightarrow 0.9x + 0.1 = 4.6$$

$$\Rightarrow x = \frac{4.5}{0.9} = 5$$

$$x = 5$$

3. Consider the strong electrolytes  $\text{Z}_m\text{X}_n$ ,  $\text{U}_m\text{Y}_p$  and  $\text{V}_m\text{X}_n$ . Limiting molar conductivity ( $\Lambda^\circ$ ) of  $\text{U}_m\text{Y}_p$  and  $\text{V}_m\text{X}_n$  are 250 and 440  $\text{S cm}^2 \text{ mol}^{-1}$ , respectively. The value of  $(m + n + p)$  is \_\_\_\_\_.

Given:

Ion	$\text{Z}^{m+}$	$\text{U}^{p+}$	$\text{V}^{m+}$	$\text{X}^{n-}$	$\text{Y}^{p-}$
$\Lambda^\circ$ ( $\text{S cm}^2 \text{ mol}^{-1}$ )	50.0	25.0	100.0	80.0	100.0

$\Lambda^\circ$  is the limiting molar conductivity of ions

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