

RATE OF CHEMICAL REACTION:

concentration of species taking part in the chemical reaction per unit time. It is the change in molar concentration of species taking part in the chemical reaction per unit time.

for the reaction $A \rightarrow B$
Rate of disappearance of A

$$= \frac{\text{Decrease in conc. of A}}{\text{Time taken}} = - \frac{\Delta[A]}{\Delta t}$$

Rate of disappearance of B

$$= \frac{\text{Increase in conc. of B}}{\text{Time taken}} = + \frac{\Delta[B]}{\Delta t}$$

NOTE The concentration of reactant decreases, so it is represented by -ve sign, while the conc. of product increases so it is represented by +ve sign.

TYPES OF RATE OF REACTION:

Average Rate

change in molar conc. of reactant and product at a given interval of time

$$\text{Avg. Rate of Reaction} \quad R \longrightarrow P \quad - \frac{\Delta[R]}{\Delta t} = + \frac{\Delta[P]}{\Delta t}$$

Instantaneous Rate

change in molar conc. of reactant and product at a given instant of time.

$$\text{Inst. Rate of Reaction} \quad - \frac{d[R]}{dt} = + \frac{d[P]}{dt}$$

RELATION BETWEEN RATE OF REACTION AND STOICHIOMETRY



$$\text{Rate of Rxn} \quad - \frac{d[A]}{dt} = - \frac{d[B]}{dt} = + \frac{d[C]}{dt} = + \frac{d[D]}{dt}$$

•) $2A \longrightarrow C + D$ It is clear from stoichiometry of reaction that the rate of disappearance of A is twice the velocity of formation of C & D.

1 Review And Reinforcement Chemistry Chemical Kinetics

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Unveiling the Magic of Words: A Review of " **1 Review And Reinforcement Chemistry Chemical Kinetics**"

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