

6.14. Density:

6.14.1. Definition:

1.000 liters of pure water (H_2O) corresponds to 1.000 kg. That is a nice convenience, but does not apply to all other liquids, different in composition from pure water (H_2O): for some liquids, a 1.000 kg of that liquid will weigh more, and for some liquids, a 1.000 kg will weigh less. Obviously, the type of liquid will have an influence, this is how got defined the specific weight, or density of a type of matter/liquid/gas:

$$\text{DENSITY} = \frac{\text{MASS}}{\text{VOLUME}} \quad \text{OR} \quad \rho = m/v$$

Density describes the ratio between the mass of a body of matter, or a liquid on the one hand and the volume it takes up, on the other hand. The symbols and unit of density:

$$\text{DENSITY } [\rho] = \frac{\text{MASS}}{\text{VOLUME}} \quad \frac{[m]}{[V]}$$

$$\text{DENSITY } [\rho] = \frac{\text{MASS}}{\text{VOLUME}} \quad \frac{[kg]}{[m^3]}$$

Density is expressed by the Greek symbol " ρ ", pronounced as "rho". The unit in the metric system is expressed in KILOGRAM/ m^3 and helps to express the mass of 1 m^3 , of a certain substance (solid, liquid, or gaseous). Density is one of the objective criteria to distinguish substances from each other. For example:

EXAMPLE 6.14.1: AvGas (Aviation Gasoline) has a density of 0,72 g/cm³.

How many kg/m³ does that correspond with?

- a. 72.000 kg/m³
- b. 720 kg/m³
- c. 72 kg/m³
- d. 7.200 kg/m³



ANSWER:

Here we are facing a conversion of 2 factors, both the mass and the volume, before we can state anything useful with regards to what was asked.

STEP 1: Let's first tackle the mass, which needs to be converted to the standard unit: g -> kg:

kilogram	hectogram <i>(rarely used unit)</i>	decagram <i>(rarely used unit)</i>	gram	decigram
0,	0	0	7	2

72 g corresponds to 0,0072 kg

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 - The Rise of Digital Reading Air Pilot Manual Maths And Physics Exams
 - Advantages of eBooks Over Traditional Books
2. Identifying Air Pilot Manual Maths And Physics Exams
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Air Pilot Manual Maths And Physics Exams
 - User-Friendly Interface
4. Exploring eBook Recommendations from Air Pilot Manual Maths And Physics Exams
 - Personalized Recommendations
 - Air Pilot Manual Maths And Physics Exams User Reviews and Ratings
 - Air Pilot Manual Maths And Physics Exams and Bestseller Lists
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 - Highlighting and Note-Taking Air Pilot Manual Maths And Physics Exams
 - Interactive Elements Air Pilot Manual Maths And Physics Exams
- 8. Staying Engaged with Air Pilot Manual Maths And Physics Exams
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Air Pilot Manual Maths And Physics Exams
- 9. Balancing eBooks and Physical Books Air Pilot Manual Maths And Physics Exams
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Air Pilot Manual Maths And Physics Exams
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Air Pilot Manual Maths And Physics Exams
 - Setting Reading Goals Air Pilot Manual Maths And Physics Exams
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Air Pilot Manual Maths And Physics Exams
 - Fact-Checking eBook Content of Air Pilot Manual Maths And Physics Exams
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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