

First PUC Supplementary Examination, April - 2018

Subject : CHEMISTRY

Time : 3 Hours 15 Min

Max. Marks : 70

Instructions :

- i) The question paper has four parts : A, B, C, D.
 - a) Part-A carries 10 marks.
 - b) Part-B carries 10 marks.
 - c) Part-C carries 15 marks.
 - d) Part-D carries 35 marks.
- iii) Write neat labelled diagrams wherever required.
- iv) Use only simple calculators. (Use of Scientific Calculator is not allowed.)
- v) Write balanced chemical equations wherever required.
- vi) Use log tables if necessary.

Part - A

I. Answer the following questions :

10 x 1 = 10

1. Give the Scientific notation of 233.508
2. What is the value of compressibility factor (z) for an ideal gas?
3. What is common ion effect?
4. Which element has highest electronegativity?
5. Find the oxidation number of 'P' in NaH_2PO_4 .
6. Give the composition of plaster of Paris.
7. What is Producer Gas?
8. Give an example for Zeolite.
9. Write the functional group of Carboxylic acids.
10. Identify the product 'A'.



Part - B

II. Answer any Five of the following questions. Each question carries two marks :

5 x 2 = 10

11. Give example for :
 - a) Element
 - b) Compound
12. Write any two differences between Ideal gas and Real gas.
13. What is Dipole Moment? Give its unit.
14. Mention any two similarities between Lithium and Magnesium (Mg).
15. What is the action of conc. H_2SO_4 on formic acid at 373K? Give equation.
16. Explain Markovnikov's rule with an example.
17. State and explain Huckel's rule of Aromaticity.
18. What is BOD? Give its significance.

Part - C

III. Answer any Five of the following questions :

5 x 3 = 15

19. What is Electron Gain Enthalpy? How does it vary along the period and down the group?
20. Explain SP hybridisation in Ethyne (C_2H_2).
21. Write the electronic configuration of Oxygen molecule using Molecular orbital theory and explain its paramagnetic property.
22. Define the terms :
 - a) Bond Enthalpy
 - b) Bond Angle
 - c) Bond Length

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