

following table.

Mass of empty beaker	60.169 g
Mass of beaker + MnCl ₂ (s)	61.093 g
Mass of beaker + MnCl ₂ (s) + H ₂ O(l)	62.617 g

(b) Calculate the mass of Cl in the sample of MnCl₂(s) remaining in the beaker.

$$\begin{aligned} \text{mass of Mn(s)} &= 61.093 \text{ g} - 60.169 \text{ g} = 1.093 \text{ g} \\ \text{mass of MnCl}_2 &= 61.093 \text{ g} - 60.169 \text{ g} = 2.504 \text{ g} \\ \text{mass Cl} &= 2.504 \text{ g} - 1.093 \text{ g} = 1.411 \text{ g} \end{aligned}$$

(c) Calculate the number of moles of Cl in the sample of MnCl₂(s) remaining in the beaker.

$$\text{moles} = \frac{\text{given mass}}{\text{molar mass}} = \frac{1.411 \text{ g}}{35.45 \text{ g/mol}} = 0.03980 \text{ mol Cl}$$

(d) The student determines that the empirical formula of the MnCl₂ is MnCl₂. Use the data to determine the

SOLVED!!

$$\begin{array}{r} \text{moles} \\ 0.0149 \text{ mol} \\ \hline 0.0149 \end{array}$$

$$\begin{array}{r} 0.03980 \text{ mol} \\ \hline 0.0149 \end{array}$$

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Princeton Review



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