



# HEATING METAL CARBONATES

**Decomposition** is the breakdown of a compound into simpler substances. Some compounds can be broken down using heat – this is **thermal decomposition**. Some compounds can be broken down using electricity – this is **electrical decomposition** (also called **electrolysis**).

Many metal carbonates break down when heated as shown: metal carbonate → metal oxide + carbon dioxide

## AIM

- 1) To find out which metal carbonates decompose on heating in a Bunsen flame.
- 2) To link the ease with which a metal carbonate decomposes to the reactivity of the metal in the compound.

Reactivity series of metals	Most reactive metal	↑	potassium magnesium zinc copper
	Least reactive metal		

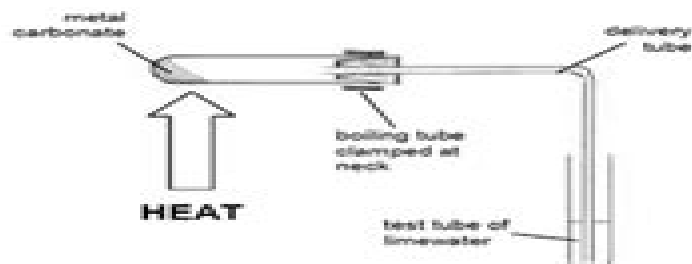
## VARIABLES

Independent variable	Dependent variable	Control variables
<input type="checkbox"/> categoric <input type="checkbox"/> continuous	<input type="checkbox"/> categoric <input type="checkbox"/> continuous	   

## METHOD

 Wear eye protection	 Be careful of hot apparatus. Remove the delivery tube out of the limewater before you stop heating.	 copper carbonate	 potassium carbonate limewater
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- 1) Place two spatula loads of one of the metal carbonates in a dry test tube clamped horizontally as shown in the diagram.
- 2) Heat the tube with a hot Bunsen flame.
- 3) Time how long it takes for the limewater to turn cloudy.
- 4) Remove the tube with the limewater before you stop heating.
- 5) Do not heat for more than 4 minutes even if it does not turn cloudy.
- 6) Repeat for each metal carbonate. Each time use a clean test tube for the metal carbonate and fresh limewater.



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**Daniela Niemeyer**



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