

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

9 3 Limiting Reagent And Percent Yield Worksheet Answers

Jessica J Manson



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Table of Contents 9 3 Limiting Reagent And Percent Yield Worksheet Answers

1. Understanding the eBook 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - The Rise of Digital Reading 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Personalized Recommendations
 - 9 3 Limiting Reagent And Percent Yield Worksheet Answers User Reviews and Ratings
 - 9 3 Limiting Reagent And Percent Yield Worksheet Answers and Bestseller Lists
5. Accessing 9 3 Limiting Reagent And Percent Yield Worksheet Answers Free and Paid eBooks

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- 9 3 Limiting Reagent And Percent Yield Worksheet Answers eBook Subscription Services
- 9 3 Limiting Reagent And Percent Yield Worksheet Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - 9 3 Limiting Reagent And Percent Yield Worksheet Answers Compatibility with Devices
 - 9 3 Limiting Reagent And Percent Yield Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Highlighting and Note-Taking 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Interactive Elements 9 3 Limiting Reagent And Percent Yield Worksheet Answers
- 8. Staying Engaged with 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 9 3 Limiting Reagent And Percent Yield Worksheet Answers
- 9. Balancing eBooks and Physical Books 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 9 3 Limiting Reagent And Percent Yield Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Setting Reading Goals 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Fact-Checking eBook Content of 9 3 Limiting Reagent And Percent Yield Worksheet Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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