

Factors affecting electronic conductance

- Nature of conductor
- No. of valence of e^-
- Temp $\uparrow$ $K \downarrow$

Factors affecting electrolytic or ionic conductance

(1) Nature

Strong electrolyte $\rightarrow$ more

(2) Nature of solvent

high polarity
high will be dissociation

(3) Temp

$T \uparrow$ $K \uparrow$ mobility $\uparrow$ $K \uparrow$

FOR NUMERICAL'S

$$R = \frac{\rho \cdot l}{A} \quad \Omega$$

$$G^* = \frac{\sigma}{l} \quad G^* = KR \quad m^{-1}$$

$$\rho = \frac{1}{K} \quad \frac{\Omega \cdot m^{-1}}{S \cdot m^{-1}}$$

$$G = \frac{1}{R} \quad \frac{S}{m}$$

$$G = \frac{1}{R} \quad \frac{S}{m}$$

$$G = \frac{1}{R} \quad \frac{S}{m}$$

molar conductivity

$$\lambda_m = \frac{K \times 1000}{M} \quad S \cdot cm^2 \cdot mol^{-1}$$

$$\lambda_m = \frac{K}{1000 M} \quad S \cdot m^2 \cdot mol^{-1}$$

Equivalent conductivity

$$\lambda_{eq} = \frac{K \times 1000}{N} \quad S \cdot cm^2 \cdot eq^{-1}$$

$$\lambda_{eq} = \frac{K}{N \times 1000} \quad S \cdot m^2 \cdot eq^{-1}$$

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Variation with conc

K always $\downarrow$ conc $\downarrow$ $V \uparrow$
(we, se) (volume)

$\lambda_m \uparrow$ conc $\downarrow$ $V \uparrow$
as conc $\downarrow$

interionic attraction $\downarrow$

Strong electrolyte

λ_m increase slowly with dilution

$$\lambda_m = \lambda_m^\infty - A \cdot c^{1/2}$$

Weak electrolytes

for we lower α at higher conc

on dilution $\alpha \uparrow$ $V \uparrow$

for we, λ_m^∞ measured by Kohlrausch

effect of dilution

dilution $\uparrow$ Volume $\uparrow$

dilution $\uparrow$ $K \downarrow$ Volume $\uparrow$

dilution $\uparrow$ $\lambda_m \uparrow$ volume $\uparrow$

magnitude of inc. in molar conductance for weak electrolyte is much larger than that for SE becoz strong electrolyte completely ionised in all conc

Kohlrausch law

independent migration of ions

$$\alpha = \frac{\lambda_m}{\lambda_m^\infty}$$

$$K_0 = \frac{C \alpha^2}{(1 - \alpha)}$$

law state that limiting molar conductivity of an electrolyte can be represented as the sum of the individual contribution of anion of the electrolyte

$$\lambda_m^\infty = \nu_+ \lambda_c + \nu_- \lambda_a$$

Electrolysis

1st law

$$W = ZQ$$

$$W = \frac{E}{F} Q$$

$$W = \frac{m}{nF} Q$$

$$W = \frac{m}{nF} It$$

$$F = 96500$$

$$\frac{W}{m} \times nF = \frac{It}{F} = \frac{Q}{F}$$

$$n \times nF = \frac{It}{F} = \frac{Q}{F}$$

$$V \times nF = n \times nF = \frac{It}{F}$$

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$$\frac{1}{N_2} = \frac{c_1}{Z_1} = \frac{c_2}{Z_2}$$

Electrochemical series

SRP $\uparrow$ Reduction cathode oxidising agent

SRP $\downarrow$ oxidation anode Reducing agent

Reactivity of metals

Reactivity $\propto \frac{1}{SRP}$ $\propto SOP$

P > C > S > M > A > Z > N > D > H

K > Ca > Na > Mg > Al > Zn > Fe > Sn > Pb

C > H > A > R > P

Cu > Hg > Ag > Au > Pt

Reactivity of non metals

Reactivity $\propto$ SRP

F > S > N > O > Cl > Br > I

Potential discharge theory

P < S < M < A < Z < N < D < H

CHARP

F < S < N < O < Cl < Br < I

ions having less discharge potential will release first

IMP FORMULA

$$E^\circ = E_c^{SRP} - E_a^{CARP}$$

$$E = E^\circ - \frac{0.059}{n} \log \frac{P}{R}$$

$$\Delta G^\circ = -nE^\circ F$$

$$E^\circ = \frac{0.06}{n} \log K$$

$$\Delta G^\circ = -2.303 nRT \log K$$

$$\Delta G^\circ = -2.303 nRT \log K$$

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