

Types of Chemical Reactions Worksheet #1

11-Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ **Synthesis**
- 2) $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$ **Single Replacement**
- 3) $\text{HNO}_3 + \text{AlCl}_3 \rightarrow \text{Al}(\text{OH})_3 + 3\text{HCl}$ **Double Replacement**
- 4) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ **Synthesis**
- *5) $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ **Double Replacement**
- 6) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ **Decomposition**
- 7) $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$ **Single Replacement**
- 8) $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$ **Double Replacement**
- 9) $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$ **Synthesis**

*10) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ **Single Replacement**

*Remember: " H_2O " can also be thought of as " HOH "

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement: $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement: $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition: $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement: $3\text{Ca}(\text{OH})_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{CaCl}_2$
- 16) Single replacement: $\text{Fe} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Cu}$ [Assume Fe^{2+}]
- 17) Decomposition: $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement: $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement: $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$ [Copper (II) is used here]

1describing Chemical Reactions Answers

Louis Theodore



1describing Chemical Reactions Answers:

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ERDA Energy Research Abstracts United States. Energy Research and Development Administration,1977

Qualitative Theory of Parabolic Equations, Part 1 T. I. Zelenyak,M. P. Vishnevskii,M. M. Lavrentiev,2011-09-06 No detailed description available for Qualitative Theory of Parabolic Equations Part 1

Dynamics of Exothermicity Brian Bowen,1996-09-15 Covering the dynamics of reactive systems and of explosions the 15 papers discuss the treatment of turbulent mixing in reactive systems acoustic interactions with combustion fields liquid atomization soot formation practical applications of combustion in waste incineration and pulse jet ignition in internal combustion engines detonations phenomena and mixing effects in explosions Includes six color plates No index Annotation copyrighted by Book News Inc Portland OR

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Sessional Papers British Columbia,1921

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