

# DESCRIBING CHEMICAL REACTIONS

## Section Review - Objectives

- Explain how to write a word equation
- Describe how to write a skeleton equation
- List the steps for writing a complete chemical equation.

## Vocabulary

- chemical equation
- skeleton equation
- catalyst
- coefficients
- balanced equation

## Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

A chemical reaction can be concisely represented by a chemical 1.

The substances that undergo a chemical change are the 2. The new substances formed in a chemical reaction are the 3. In accordance with the law of conservation of 4, a chemical equation must be balanced.

When balancing an equation, you place 5 in front of reactants and products so that the same number of atoms of each 6 are on each side of the equation. An equation must never be balanced by changing the 7 in the chemical formula of a substance.

Special symbols are used to show the physical state of a substance in a reaction. The symbol for a liquid is 8; for a solid, 9; and for a gas, 10. A substance dissolved in water is designated 11. If a 12 is used to increase the rate of a chemical reaction, its formula is written above the arrow.

1. \_\_\_\_\_
2. \_\_\_\_\_
3. \_\_\_\_\_
4. \_\_\_\_\_
5. \_\_\_\_\_
6. \_\_\_\_\_
7. \_\_\_\_\_
8. \_\_\_\_\_
9. \_\_\_\_\_
10. \_\_\_\_\_
11. \_\_\_\_\_
12. \_\_\_\_\_

## Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- \_\_\_\_\_ 13. In an equation, a substance is shown to be in the gaseous state by placing an upward-pointing arrow after its formula.
- \_\_\_\_\_ 14. The symbol  $\Delta$  placed over the arrow in an equation means that heat is supplied to the reaction.
- \_\_\_\_\_ 15. Atoms are destroyed in a chemical reaction.
- \_\_\_\_\_ 16. A skeleton equation is not a balanced equation.

## Part D Questions and Problems

24. Write a balanced equation for each of these chemical reactions. I

a. Aluminum reacts with aqueous hydrochloric acid to form hydrogen gas and aqueous aluminum chloride.

b. Acetylene gas ( $C_2H_2$ ) burns in a welding torch with oxygen to form carbon dioxide gas and water vapor.

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