

Balancing Equations Worksheet II

Count the number of atoms, then use coefficients to balance each equation.

- 1) $\underline{\hspace{1cm}} \text{AlBr}_3 + \underline{\hspace{1cm}} \text{K} \rightarrow \underline{\hspace{1cm}} \text{KBr} + \underline{\hspace{1cm}} \text{Al}$
- 2) $\underline{\hspace{1cm}} \text{FeO} + \underline{\hspace{1cm}} \text{PdF}_2 \rightarrow \underline{\hspace{1cm}} \text{FeF}_2 + \underline{\hspace{1cm}} \text{PdO}$
- 3) $\underline{\hspace{1cm}} \text{P}_4 + \underline{\hspace{1cm}} \text{Br}_2 \rightarrow \underline{\hspace{1cm}} \text{PBr}_3$
- 4) $\underline{\hspace{1cm}} \text{LiCl} + \underline{\hspace{1cm}} \text{Br}_2 \rightarrow \underline{\hspace{1cm}} \text{LiBr} + \underline{\hspace{1cm}} \text{Cl}_2$
- 5) $\underline{\hspace{1cm}} \text{PbBr}_2 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{HBr} + \underline{\hspace{1cm}} \text{PbCl}_2$
- 6) $\underline{\hspace{1cm}} \text{CoBr}_3 + \underline{\hspace{1cm}} \text{CaSO}_4 \rightarrow \underline{\hspace{1cm}} \text{CaBr}_2 + \underline{\hspace{1cm}} \text{Co}_2(\text{SO}_4)_3$
- 7) $\underline{\hspace{1cm}} \text{Na}_3\text{P} + \underline{\hspace{1cm}} \text{CaF}_2 \rightarrow \underline{\hspace{1cm}} \text{NaF} + \underline{\hspace{1cm}} \text{Ca}_3\text{P}_2$
- 8) $\underline{\hspace{1cm}} \text{Mn} + \underline{\hspace{1cm}} \text{HI} \rightarrow \underline{\hspace{1cm}} \text{H}_2 + \underline{\hspace{1cm}} \text{MnI}_3$
- 9) $\underline{\hspace{1cm}} \text{Li}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{NaBr} \rightarrow \underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{LiBr}$
- 10) $\underline{\hspace{1cm}} \text{CaF}_2 + \underline{\hspace{1cm}} \text{Li}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{CaSO}_4 + \underline{\hspace{1cm}} \text{LiF}$
- 11) $\underline{\hspace{1cm}} \text{HBr} + \underline{\hspace{1cm}} \text{Mg}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{MgBr}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 12) $\underline{\hspace{1cm}} \text{LiNO}_3 + \underline{\hspace{1cm}} \text{CaBr}_2 \rightarrow \underline{\hspace{1cm}} \text{Ca}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{LiBr}$
- 13) $\underline{\hspace{1cm}} \text{AgNO}_3 + \underline{\hspace{1cm}} \text{Li} \rightarrow \underline{\hspace{1cm}} \text{LiNO}_3 + \underline{\hspace{1cm}} \text{Ag}$
- 14) $\underline{\hspace{1cm}} \text{Si}(\text{OH})_4 + \underline{\hspace{1cm}} \text{NaBr} \rightarrow \underline{\hspace{1cm}} \text{SiBr}_4 + \underline{\hspace{1cm}} \text{NaOH}$
- 15) $\underline{\hspace{1cm}} \text{NaCN} + \underline{\hspace{1cm}} \text{CuCO}_3 \rightarrow \underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{Cu}(\text{CN})_2$

Challenge: $\underline{\hspace{1cm}} \text{C}_2\text{H}_6 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$

1chemical Reactions Practice Problems Key

**Patrick E. McMahon, Bohdan B.
Khomtchouk, Claes Wahlestedt**



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Table of Contents 1chemical Reactions Practice Problems Key

1. Understanding the eBook 1chemical Reactions Practice Problems Key
 - The Rise of Digital Reading 1chemical Reactions Practice Problems Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 1chemical Reactions Practice Problems Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 1chemical Reactions Practice Problems Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 1chemical Reactions Practice Problems Key
 - Personalized Recommendations
 - 1chemical Reactions Practice Problems Key User Reviews and Ratings
 - 1chemical Reactions Practice Problems Key and Bestseller Lists
5. Accessing 1chemical Reactions Practice Problems Key Free and Paid eBooks
 - 1chemical Reactions Practice Problems Key Public Domain eBooks
 - 1chemical Reactions Practice Problems Key eBook Subscription Services
 - 1chemical Reactions Practice Problems Key Budget-Friendly Options

6. Navigating 1chemical Reactions Practice Problems Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 1chemical Reactions Practice Problems Key Compatibility with Devices
 - 1chemical Reactions Practice Problems Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 1chemical Reactions Practice Problems Key
 - Highlighting and Note-Taking 1chemical Reactions Practice Problems Key
 - Interactive Elements 1chemical Reactions Practice Problems Key
8. Staying Engaged with 1chemical Reactions Practice Problems Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 1chemical Reactions Practice Problems Key
9. Balancing eBooks and Physical Books 1chemical Reactions Practice Problems Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1chemical Reactions Practice Problems Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 1chemical Reactions Practice Problems Key
 - Setting Reading Goals 1chemical Reactions Practice Problems Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 1chemical Reactions Practice Problems Key
 - Fact-Checking eBook Content of 1chemical Reactions Practice Problems Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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