# Balancing Equations Worksheet II

Count the number of atoms, then use coefficients to balance each equation.

AlBr<sub>3</sub> + \_\_\_ K → \_\_\_ KBr + \_\_\_ Al

2) \_\_\_\_ FeO + \_\_\_ PdF₂ → \_\_\_ FeF₂ + \_\_\_ PdO

3) \_\_\_\_ P<sub>4</sub> + \_\_\_\_ Br<sub>2</sub> → \_\_\_\_ PBr<sub>3</sub>

4) \_\_\_\_ LiCl + \_\_\_ Br<sub>2</sub> → \_\_\_ LiBr + \_\_\_ Cl<sub>2</sub>

5) \_\_\_\_ PbBr<sub>2</sub> + \_\_\_ HCl → \_\_\_ HBr + \_\_\_ PbCl<sub>2</sub>

6) \_\_\_\_ CoBr<sub>3</sub> + \_\_\_\_ CaSO<sub>4</sub> → \_\_\_\_ CaBr<sub>2</sub> + \_\_\_\_ Co<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>

7) \_\_\_\_ Na<sub>3</sub>P + \_\_\_ CaF<sub>2</sub> → \_\_\_ NaF + \_\_\_ Ca<sub>3</sub>P<sub>2</sub>

8) \_\_\_\_ Mn + \_\_\_ HI → \_\_\_ H₂ + \_\_\_ MnI₃

9) \_\_\_\_ Li<sub>3</sub>PO<sub>4</sub> + \_\_\_\_ NaBr → \_\_\_\_ Na<sub>3</sub>PO<sub>4</sub> + \_\_\_\_ LiBr

10) \_\_\_\_ CaF<sub>2</sub> + \_\_\_\_ Li<sub>2</sub>SO<sub>4</sub> → \_\_\_\_ CaSO<sub>4</sub> + \_\_\_\_ LiF

11) \_\_\_\_ HBr + \_\_\_ Mg(OH)<sub>2</sub> → \_\_\_ MgBr<sub>2</sub> + \_\_\_ H<sub>2</sub>O

12) \_\_\_\_ LiNO<sub>3</sub> + \_\_\_\_ CaBr<sub>2</sub> → \_\_\_\_ Ca(NO<sub>3</sub>)<sub>2</sub> + \_\_\_\_ LiBr

13) \_\_\_\_ AgNO<sub>3</sub> + \_\_\_ Li → \_\_\_ LiNO<sub>3</sub> + \_\_\_ Ag

14) \_\_\_\_ Si(OH)₄ + \_\_\_ NaBr → \_\_\_ SiBr₄ + \_\_\_ NaOH

15) \_\_\_\_ NaCN + \_\_\_ CuCO<sub>3</sub> → \_\_\_ Na<sub>2</sub>CO<sub>3</sub> + \_\_\_ Cu(CN)<sub>2</sub>

# **1chemical Reactions Practice Problems Answers**

**Donald A. McQuarrie, Stanley Gill** 

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