

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?



$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mole Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$= \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{1 \text{ mole AgNO}_3} \times \frac{2 \text{ mol AgNO}_3}{1 \text{ mole Cu}} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} = \underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole Cl}_2}{2 \text{ mole NaCl}} \times \frac{71 \text{ g Cl}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mol Cl}_2}{1 \text{ mole Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{0.21 \text{ moles H}_2}$$

Practice Problems Chemistry Answers

Kenneth W. Raymond



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Student Solutions Manual Darrell Ebbing, 2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions

General Organic and Biological Chemistry Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features

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Partial Solutions Manual Darrell D. Ebbing, David Bookin, 1998

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Table of Contents Practice Problems Chemistry Answers

1. Understanding the eBook Practice Problems Chemistry Answers
 - The Rise of Digital Reading Practice Problems Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Problems Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Problems Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Problems Chemistry Answers
 - Personalized Recommendations
 - Practice Problems Chemistry Answers User Reviews and Ratings
 - Practice Problems Chemistry Answers and Bestseller Lists
5. Accessing Practice Problems Chemistry Answers Free and Paid eBooks
 - Practice Problems Chemistry Answers Public Domain eBooks
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 - Practice Problems Chemistry Answers Budget-Friendly Options
6. Navigating Practice Problems Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Problems Chemistry Answers Compatibility with Devices
 - Practice Problems Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Problems Chemistry Answers
 - Highlighting and Note-Taking Practice Problems Chemistry Answers
 - Interactive Elements Practice Problems Chemistry Answers
8. Staying Engaged with Practice Problems Chemistry Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Practice Problems Chemistry Answers
- 9. Balancing eBooks and Physical Books Practice Problems Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Problems Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Problems Chemistry Answers
 - Setting Reading Goals Practice Problems Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Problems Chemistry Answers
 - Fact-Checking eBook Content of Practice Problems Chemistry Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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