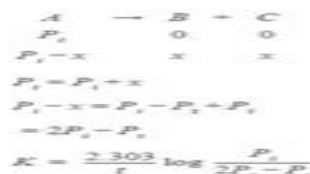


D.

$$k = \frac{2.303}{t} \times \frac{P_1}{(2P_1 - P_2)}$$

Answer: A

Solution:



Question7

$r = k[A]$ for a reaction, 50% of A is decomposed in 120 minutes. The time taken for 90% decomposition of A is ____ minutes.

[31-Jan-2024 Shift 2]

Answer: 399

Solution:

$$r = k[A]$$

So, order of reaction = 1

$$t_{1/2} = 120 \text{ min}$$

For 90% completion of reaction

$$\Rightarrow k = \frac{2.303}{t} \log \left(\frac{a}{a-x} \right)$$

$$\Rightarrow \frac{0.693}{t_{1/2}} = \frac{2.303}{t} \log \frac{100}{10}$$

$$\therefore t = 399 \text{ min.}$$

Question8

The ratio of $^{14}\text{C}/^{12}\text{C}$ in a piece of wood is $1/8$ part that of atmosphere. If half life of ^{14}C is 5730 years, the age of wood sample is years.

[1-Feb-2024 Shift 1]

Answer: 17190

Solution:

$$\lambda t = \ln \frac{(^{14}\text{C}/^{12}\text{C})_{\text{atmosphere}}}{(^{14}\text{C}/^{12}\text{C})_{\text{wood sample}}}$$

As per the question,

$$\frac{(^{14}\text{C}/^{12}\text{C})_{\text{wood}}}{(^{14}\text{C}/^{12}\text{C})_{\text{atmosphere}}} = \frac{1}{8}$$

$$\text{So, } \lambda t = \ln 8$$

$$\frac{\ln 2}{t_{1/2}} t = \ln 8$$

$$t = 3 \times t_{1/2} = 17190 \text{ years}$$

Question9

The following data were obtained during the first order thermal decomposition of a gas A at constant volume:
 $\text{A(g)} \longrightarrow 2\text{B(g)} + \text{C(g)}$

S.No	Time/s	Total pressure/(atm)
1.	0	0.1
2.	400	0.15

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A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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