

Use the activity series of elements to predict which reactions take place.

Metal	Metal Ions			
	Copper (Cu)	High (H)	Low (L)	Ag (Ag)
Copper (Cu)	<b>NO</b>	<b>NO</b> More active	<b>YES</b> More active	<b>NO</b>
Magnesium (Mg)	<b>YES</b>	<b>NO</b>	<b>NO</b>	<b>NO</b>
Zinc (Zn)	<b>YES</b>	<b>NO</b> More active	<b>NO</b>	<b>NO</b>
Silver (Ag)	<b>NO</b> More active	<b>NO</b> More active	<b>YES</b> More active	<b>NO</b>

### Questions

1. Write your cell as a spontaneous redox reaction, balanced using the half-cell method.



2. Write reaction with the most activity? **Magnesium**

3. Write reaction with the least activity? **Ag**

4. List the four metals in order, decreasing from most active to least active, with the most activity on the left and the least activity on the right. The metal that reacts with the least activity is on the left and the metal that reacts with the most activity is on the right.



5. Write the net ionic equation for the reaction. Write the net ionic equation for the reaction of magnesium metal with copper ions. Is the reaction spontaneous?



6. Write the net ionic equation for the reaction of silver ions with copper metal. Is the reaction spontaneous?



7. Write the net ionic equation for the reaction of silver ions with copper metal. Is the reaction spontaneous?

# An Activity Series Lab Answers

**United States. Congress. House.  
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